**7th Grade Science Lesson Plans**

October 28 – November 1

Modeling Conservation of Mass

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|  | **MONDAY** | **TUESDAY** | **WEDNESDAY** | **THURSDAY** | **FRIDAY** |
| **CONTENT**  **OBJECTIVE:** | Students will be able to demonstrate ***analysis*** of conservation of matter by ***creating and explaining*** models of chemical reaction equations on a type 2 writing. | Students will be able to demonstrate ***application*** of chemical reaction equations by ***balancing*** an equation on an exit ticket with a score of 2 or higher. | Students will be able to demonstrate ***synthesis*** of conservation of matter in an open and closed system by ***designing*** an experiment to test and compare. | ½ Day | Students will be able to demonstrate ***analysis*** of closed and open systems by ***explaining*** why a closed system demonstrates the law of conservation of matter using a type 2. |
| **LANGUAGE OBJECTIVE:** | I can create a written and illustrated model of a balanced chemical reaction equation. | Students will listen to the steps for balancing and respond using accountable talk. | Students will orally discuss the components needed for an experiment. |  | Students will write to explain using a type 2 writing. |
| **VOCABULARY:** | Law of Conservation of Mass |  |  |  | Closed System, Open System |
| **NGSS:** | **MS-PS1-5:**  Develop and use a model to describe how the total number of atoms does not change in a chemical reaction, and thus mass is conserved. | **MS-PS1-5:**  Develop and use a model to describe how the total number of atoms does not change in a chemical reaction, and thus mass is conserved. | **MS-PS1-5:**  Develop and use a model to describe how the total number of atoms does not change in a chemical reaction, and thus mass is conserved. |  | **MS-PS1-5:**  Develop and use a model to describe how the total number of atoms does not change in a chemical reaction, and thus mass is conserved. |